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Eisaman

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- (54) **ENERGY EFFICIENT METHOD FOR STRIPPING CO₂ FROM SEAWATER**
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 See application file for complete search history.

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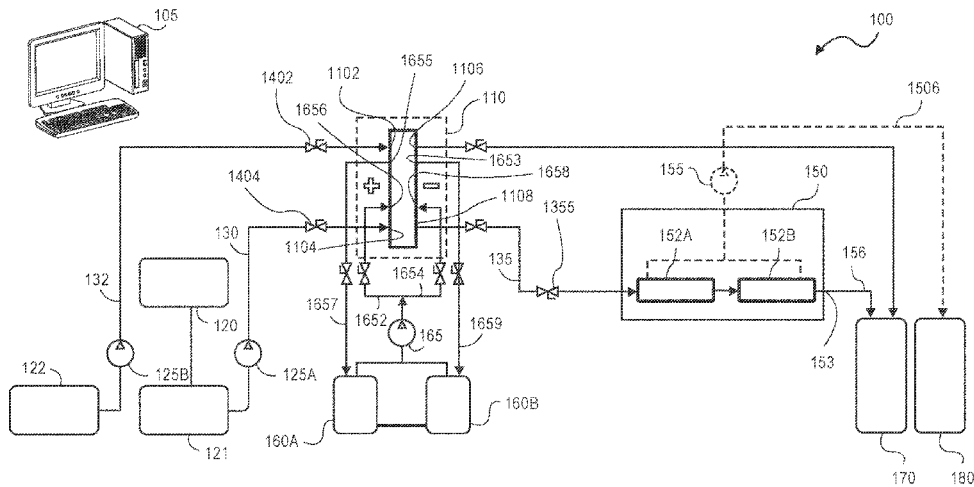
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(57) **ABSTRACT**

A method including increasing modifying a volume of seawater that holds an amount of dissolved inorganic carbon; acidifying the amount of seawater; and collecting an amount of carbon dioxide from the acidified seawater. A system including an electro dialysis unit including an acidified solution compartment, a basified solution compartment, a membrane and an acidified solution output compartment; a vessel coupled to an inlet of the acidified solution compartment and operable to contain a modified volume of seawater therein; and a desorption unit coupled to the acidified compartment output, the desorption unit operable to receive carbon dioxide gas from a solution from the acidified output compartment.

16 Claims, 3 Drawing Sheets



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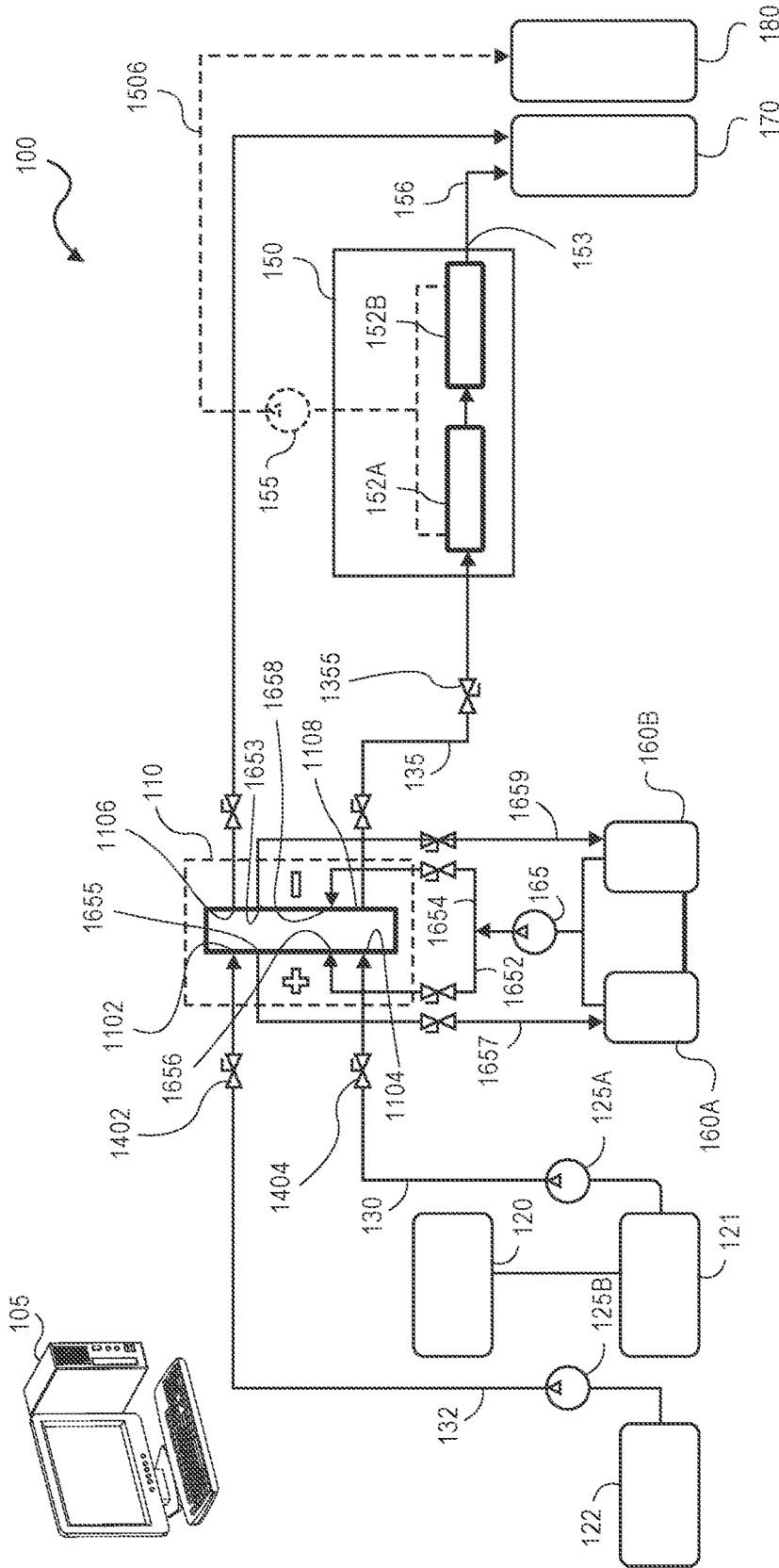


FIG. 1

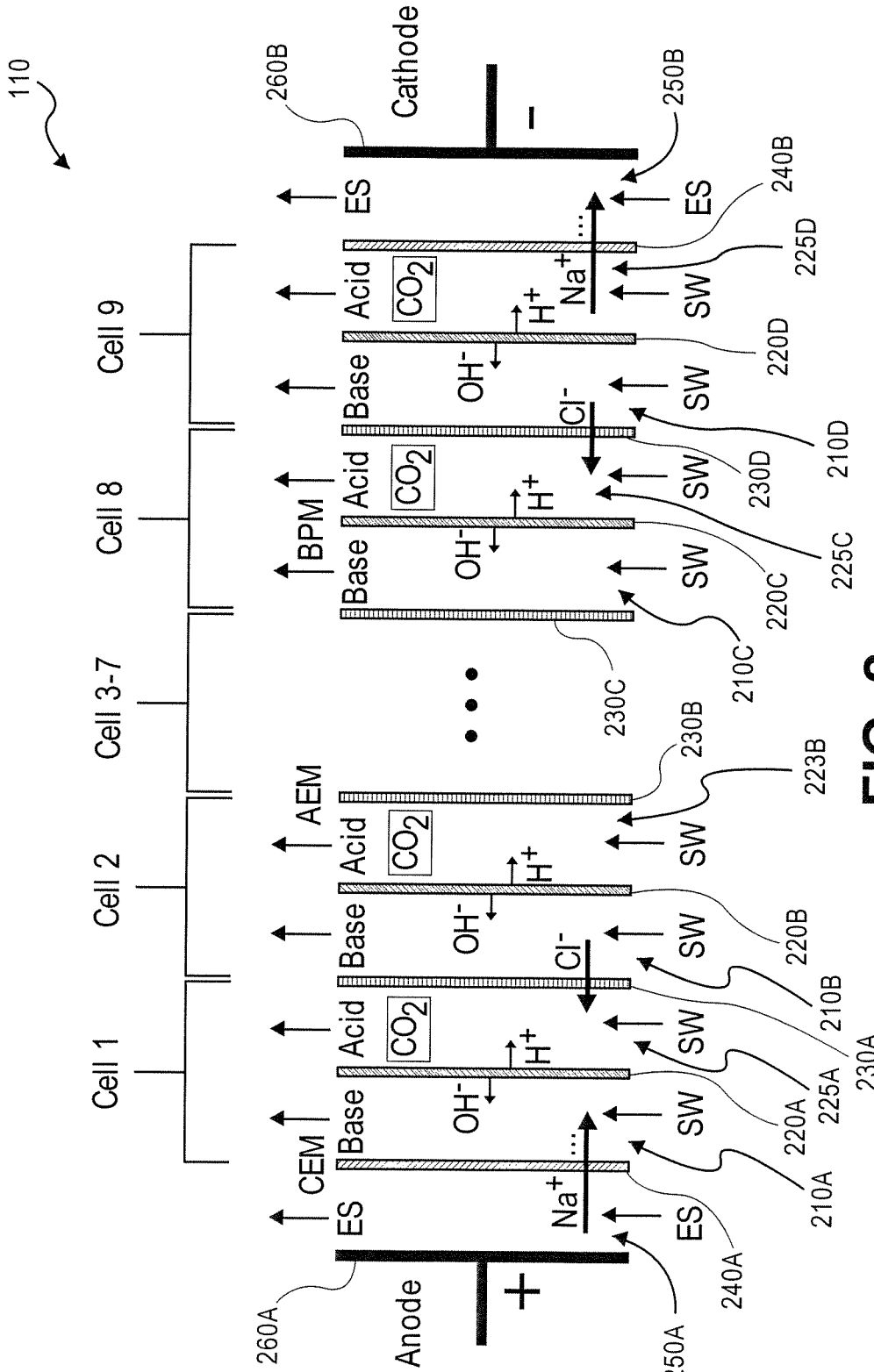


FIG. 2

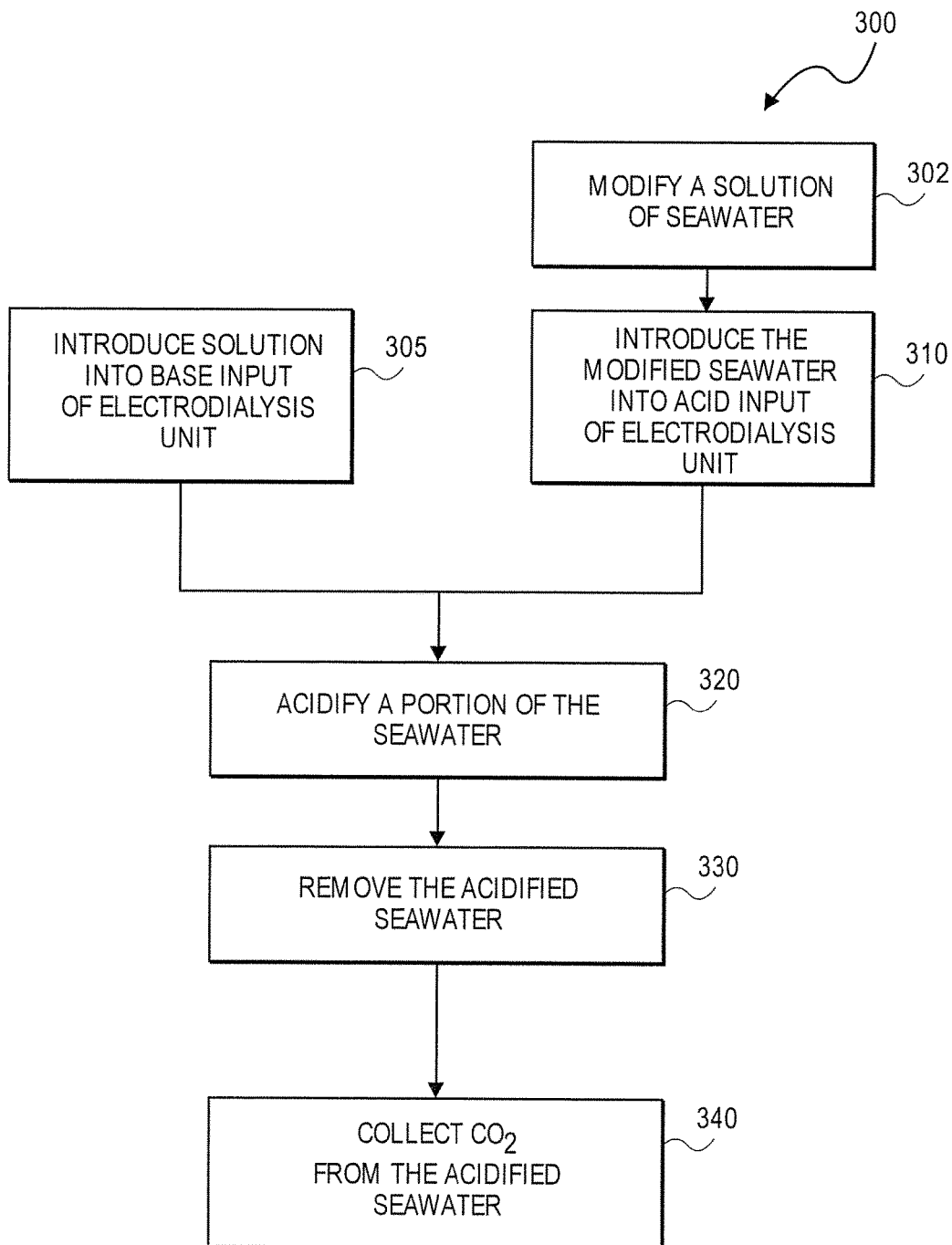


FIG. 3

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ENERGY EFFICIENT METHOD FOR
STRIPPING CO₂ FROM SEAWATER

FIELD

This disclosure relates generally to the field of carbon dioxide separation and collection.

BACKGROUND

The separation of carbon dioxide (CO₂) from a mixed-gas source may be accomplished by a capture and regeneration process. More specifically, the process generally includes a selective capture of CO₂, by, for example, contacting a mixed-gas source with a solid or liquid adsorber or absorber followed by a generation or desorption of CO₂ from the adsorber or absorber. One technique describes the use of bipolar membrane electro dialysis for CO₂ removal from potassium carbonate and bicarbonate solutions.

For capture/regeneration systems, a volume of gas that is processed is generally inversely related to a concentration of CO₂ in the mixed-gas source, adding significant challenges to the separation of CO₂ from dilute sources such as the atmosphere. CO₂ in the atmosphere, however, establishes equilibrium with the total dissolved inorganic carbon in the oceans, which is largely in the form of bicarbonate ions (HCO₃⁻) at an ocean pH of 8.1-8.3. Therefore, a method for extracting CO₂ from the ocean would effectively enable the separation of CO₂ from atmosphere without the need to process large volumes of air.

SUMMARY

In one embodiment, a method of separating carbon dioxide from seawater is described, the method including modifying a volume of seawater that holds a given amount (e.g., a mass) of dissolved inorganic carbon by, for example, increasing a concentration of dissolved inorganic carbon in the seawater and then acidifying the modified seawater followed by collecting an amount of carbon dioxide from the acidified seawater. Dissolved inorganic carbon or "DIC" as used herein includes dissolved CO₂ gas, bicarbonate ions and carbonate ions. One way the concentration of DIC in an amount of seawater is concentrated is by evaporating water molecules from the seawater (e.g., heating the seawater). By increasing a concentration of DIC in an amount of seawater, when such bicarbonate ions and carbonate ions react with hydrogen ion (upon acidification of the seawater), a partial pressure of carbon dioxide in the acidified seawater is increased. Increasing the partial pressure of carbon dioxide in seawater means less energy is required to remove (strip) the carbon dioxide from the acidified seawater.

In another embodiment, a system is described. A system includes an electro dialysis unit comprising an acidified solution compartment, a basified solution compartment, a membrane and an acidified solution output compartment and a vessel coupled to an inlet of the acidified solution compartment and operable to contain a modified volume of seawater therein. A suitable vessel is, for example, a solar pond where an amount (e.g., a volume) of seawater may be modified by evaporation of water molecules contained therein. The system also includes a desorption unit coupled to the acidified compartment output, the desorption unit operable to receive carbon dioxide gas from a solution from the acidified output compartment. In one embodiment, the desorption unit is a vessel that may receive carbon dioxide

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gas by such gas escaping from a solution from the acidified solution compartment of the electro dialysis in the absence of a vacuum.

BRIEF DESCRIPTION OF THE DRAWINGS

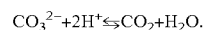
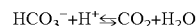
FIG. 1 shows a schematic side view of a system operable to extract or otherwise remove carbon dioxide (CO₂) from seawater.

FIG. 2 shows a cross-sectional side view of an embodiment of a bipolar membrane electro dialysis unit in the system of FIG. 1.

FIG. 3 presents a flow chart of a representative method of extracting CO₂ from seawater.

DETAILED DESCRIPTION

FIG. 1 presents a schematic representation of a system operable to extract or otherwise remove CO₂ from seawater. Referring to FIG. 1, system 100 includes electro dialysis unit 110, representatively a bipolar membrane electro dialysis (BPMED) unit that, in one embodiment, is a multi-cell membrane stack that in the presence of an electric field converts seawater (a salt solution) into two separate output streams: acidified seawater and basified seawater. The reactions indicative of CO₂ formation are:



CO₂ gas is then collected from the portion (volume) of the acidified seawater solution output from BPMED unit 110. In one embodiment, the acidified seawater solution is subjected to a desorption process wherein, for example, CO₂ and oxygen (O₂) and nitrogen (N₂) gases are vacuum stripped from the solution. To separate CO₂ from O₂ and N₂, the gas mixture can be introduced into a solution including hydroxide ions (OH⁻). The CO₂ selectively reacts with hydroxide ions in the solution to the exclusion of N₂ and O₂ to form bicarbonate (HCO₃⁻) and/or carbonate (CO₃²⁻) ions in an acidified solution. The solution containing HCO₃⁻ and/or CO₃²⁻ may be passed through a BPMED unit to acidify and desorb CO₂ that is then collected. In another embodiment, O₂ and N₂ are vacuum stripped from the modified seawater prior to acidification.

Referring to FIG. 1, system 100 includes vessel 120 having a volume that is operable to contain a volume of seawater solution (on the order of pH 8). Common seawater introduced into vessel 120 includes an amount (e.g., a mass) of DIC (e.g., dissolved CO₂ gas, bicarbonate ions (HCO₃⁻) and/or carbonate ions (CO₃²⁻)) at a concentration generally on the order of 2.5 millimolar (mM). When the concentration of DIC in common seawater is converted to carbon dioxide in an electro dialysis unit such as BPMED unit 110 (e.g., via an acid addition reaction), a partial pressure of CO₂ in the thus created acidified seawater is on the order of 0.08 atmospheres (atm) ("common acidified seawater"). Generally speaking, to thereafter separate CO₂ at a partial pressure on the order of 0.08 atmospheres from the common acidified seawater requires a reduced pressure above the solution (i.e., a pressure less than 0.08 atmospheres). Creating a pressure less than 0.08 atmospheres (i.e., creating a vacuum environment) above the solution requires energy, such as energy to power a vacuum stripper.

In one embodiment, to reduce the energy required to remove or strip CO₂ from seawater, once in vessel 120, a volume of common seawater is modified to increase a

concentration of DIC prior to an acid addition reaction to produce CO₂ (“modified seawater”). In one embodiment, once in vessel **120**, a concentration of DIC in the common seawater is increased by evaporating a portion of the water molecules therein. One technique for evaporating water molecules in the common seawater is by exposing the common seawater in vessel **120** to heat. To achieve energy savings, in one embodiment, the common seawater is exposed to low grade heat. An example of an implementation of low grade heat is where vessel **120** is a solar pond (e.g., a pond with a large surface area and shallow (low) depth). The heat for evaporation is provided by sunlight (solar thermal energy).

In one embodiment, a concentration of DIC in an amount of common seawater is increased by, for example, the evaporation of water molecules to a concentration such that, following an acid addition reaction to produce CO₂, a partial pressure of CO₂ in the acidified seawater is greater than 0.08 atmospheres and, in another embodiment, is greater than one (1) atmosphere. A partial pressure of CO₂ in acidified seawater greater than 1 atmosphere will allow for the stripping or release of CO₂ from acidified seawater in the absence of a vacuum above the acidified seawater and the associated energy requirement to create the vacuum.

Referring to FIG. 1, modified seawater from vessel **120**, such as a solar pond, in one embodiment, is optionally transferred to cooling vessel **121** where the modified seawater is cooled by, for example, ambient cooling. Representatively, cooling vessel **121** is a vessel that does not promote collection of solar thermal energy. Therefore, the modified seawater is pumped using pump **125A** through conduit **130** toward BPMED unit **110** to feed acid input **1104** of BPMED unit **110**. Separately, system **100** includes vessel **122** having contained therein, for example, an amount (e.g., a volume) of common seawater that is not modified to increase a concentration of DIC. Common seawater from vessel **122** is pumped using pump **125B** through conduit **132** toward BPMED unit **110** to feed base input **1102** of BPMED unit **110**. Base input **1102** feeds basified solution compartments within BPMED unit **110** and acid input **1104** feeds acidified solution compartments within the unit. In one embodiment, system **100** includes valve **1402** in conduit **132** to control a flow of common seawater from input tank **122** into base input **1102** and valve **1404** in conduit **130** to control a flow of modified seawater from input tank **122** into acid input **1104**.

FIG. 2 shows a cross-section of an embodiment of BPMED unit **110**. In this embodiment, the unit representatively consists of nine cells in series, with each cell consisting of: a basified solution compartment (compartments **210A**, **210B**, **210C** and **210D** illustrated); a bipolar membrane (BPM) such as Neosepta BP-1E, commercially available from Ameridia Corp. (BPM **220A**, **220B**, **220C** and **220D** illustrated); an acidified solution compartment; and an anion exchange membrane (AEM), such as Neosepta ACS, commercially available from Ameridia Corp. (AEM **230A**, **230B**, **230C** and **230D** illustrated). At each end of the membrane stack, a cation exchange membrane (CEM) such as AEM, Neosepta CMX-S, commercially available from Ameridia Corp. is used to separate the membrane stack from electrode compartment (CEM **240A** and CEM **240B** are illustrated separating the membrane stack from electrode compartment **250A** and electrode compartment **250B**, respectively). Broadly speaking, under an applied voltage, water dissociation inside the BPM and the ion-selective membranes comprising a BPM will result in the transport of H⁺ ions from one side of a BPM, and OH⁻ ions from the

opposite side. AEMs/CEMs, as their names suggest, allow the transport of negatively/positively charged ions through the membrane. The properties of these membranes such as electrical resistance, burst strength, and thickness are provided by the manufacturer (e.g., Neosepta ACS and CMX-S are monovalent-anion and monovalent-cation permselective membranes, respectively). In one embodiment, BPMED unit **110** includes electrodes **260A** and **260B** of, for example, titanium with an iridium-ruthenium based coating manufactured by De Nora Tech Inc. The solution compartments between adjacent membranes (basified solution compartments **210A-210D** and acidified solution compartments **225A-225D**) are filled with polyethylene mesh spacers (e.g., 762 μm thick polyethylene mesh spacers), and these compartments are sealed against leaks using axial pressure and 794 mm thick EPDM rubber gaskets. Each membrane has an active area of 180 cm².

In addition to the introducing seawater to BPMED unit **110**, in system **100**, an electrode solution is pumped through anode and cathode compartments of the BPMED unit, respectively. In one embodiment, a suitable electrode solution is a 0.1 molar H₂SO₄/0.25 molar Na₂SO₄ solution. FIG. 1 shows electrode solution tanks **160A** and **160B** each having a volume operable to contain an electrode solution or solutions. The electrode solution is pumped by pump **165** into conduit **1652** and conduit **1654** that are connected to electrode solution input port **1656** and electrode solution input port **1658** of BPMED unit **110** associated with an anode and a cathode, respectively. The electrode solution flows through an anode compartment and a cathode compartment of BPMED unit **110** and exits BPMED unit **110** through electrode solution output port **1655** and electrode solution output port **1653**, respectively. The electrode solution is returned to electrode solution tank **160A** and electrode solution tank **160B** through conduit **1657** and conduit **1659**, respectively. Any gases present in the anode and cathode output electrode solutions may be separately vented before recombining the solutions in the electrode solution tank.

In one embodiment, the modified seawater and common seawater are pumped into the respective acid input **1104** and base input **1102** of BPMED unit **110** and the electrode solution is pumped through anode and cathode compartments of the unit. A power supply is connected to electrodes of BPMED unit **110** and a desired current is introduced across the membrane stack to promote ion transport to produce an acidified seawater solution and a basified seawater solution.

In the embodiment shown in FIG. 1, BPMED unit **110** includes base solution output **1106** through which a basified solution is discharged and acid solution output **1108** through which an acidified solution is discharged (acidified seawater solution). The acidified solution contains dissolved CO₂ and may contain N₂ and O₂. The acidified solution discharged from acid solution output **1108** is delivered to conduit **135** connected to the output to feed CO₂ desorption unit **150**.

To extract or otherwise remove CO₂ from seawater, in one embodiment, the acidified seawater from BPMED unit **110** is directed to desorption unit **150**. In one embodiment, where a concentration of modified seawater fed to BPMED unit is such that a partial pressure of CO₂ in the acidified seawater discharged through output **1108** is greater than 1 atm, desorption unit **150** is a vessel or tank that is evacuated of air prior to receiving the acidified seawater. Upon introduction to desorption unit **150**, CO₂ in the acidified seawater will escape from solution into the tank by way of escaping to a lower pressure. The solution in desorption unit is

released at output **153** into waste tank **170** by way of conduit **156**. The CO₂ released from solution is captured in desorption unit **150**.

In another embodiment, a partial pressure in the acidified seawater is 1 atm or less (but greater than 0.08 atm). In this embodiment, stripping of CO₂ from the acidified seawater is necessary though at a lower energy requirement than prior art systems. Representatively, desorption unit **150** includes one or more membrane contactors (contactor **152A** and contactor **152B** illustrated). A suitable membrane contactor is a Liqui-Cel®X50 fibre type 2.5×8 membrane contactor commercially available from Membrana of Charlotte, N.C. Each membrane contactor has an inlet and an outlet for vacuum and an inlet and an outlet for the liquid solution to allow vacuum stripping of CO₂ from the acidified seawater solution. FIG. 1 shows vacuum pump **155** to pull a vacuum and remove CO₂ in gaseous form from the acidified seawater. CO₂ gas is removed from (exits) desorption unit **150** and transferred through conduit **175** to collection tank **180**. FIG. 1 also shows the basified output solution from BPMED unit **110** directed to waste tank **170**.

In one embodiment, an operation of system **100** described above may be controlled by a controller. FIG. 1 shows controller **105** that may be connected to through wires or wirelessly to various units of system **100** such as pump **125A** to transfer modified seawater into BPMED unit **110** from input vessel **120**; pump **125B** to transfer a solution from vessel **122** into BPMED unit **110**; pump **165** to transfer electrode solution into BPMED unit **110** from electrode solution tanks **160A** and **160B**; and optional vacuum pump **155** to extract CO₂ gas from desorption unit **150**. Variable pumps (IDNM 3534 motor and VS1MX Microdrive, Baldor Electric Company) can be used to control the flow rate and pressure of modified seawater and electrode solution. In addition to the described pumps, in one embodiment, controller **105** is connected to various valves to control the respective flow of liquids and gases through the system. Representatively, controller **105** is connected to valves **1402** and **1404** to control the flow of a solution from input tank **122** into base input **1102** of BPMED **110** and the flow of modified seawater from input tank **120** into acid input **1104** of BPMED unit **110**. Other valves that are representatively controlled by controller **105** include valves positioned in conduit at the output side of BPMED unit **110** including, but not limited to, valve **1355** to control the flow of the acidified seawater output from BPMED unit **110** to desorption unit **150**.

In one embodiment, controller **105** contains machine-readable program instructions as a form of non-transitory media. In one embodiment, the program instructions perform a method of extracting and collecting CO₂ from seawater. FIG. 3 presents a flow chart of a representative method. Referring to FIG. 3, method **300** describes modifying a volume of seawater (block **302**). In one embodiment, a volume of seawater may be modified to increase a concentration of DIC. Method **300** also describes introducing the modified seawater into an acid input of an electro dialysis unit such as BPMED unit **110** in FIG. 1 (block **310**) and introducing a solution such as unmodified seawater into a base input of the unit. The program instructions associated with controller **105** direct, for example, the opening of valves **1402** and **1403** and the operation of pump **125A** to transfer the modified seawater from input tank **120** and pump **125B** to transfer a solution from tank **122**. The program instructions similarly control the introduction of electrode solution into the electro dialysis unit and a control of current through the unit to create an electrical potential

across the membrane stack and acidify seawater introduced into an acidified solution compartment of the electro dialysis unit (seawater introduced into the acid input) (block **320**). The acidified seawater being removed from the electro dialysis unit is directed to a desorption unit (e.g., desorption unit **150**, FIG. 1) where CO₂ is separated and subsequently collected.

In one embodiment, controller **105** also regulates and monitors the system. Such regulation and monitoring may be accomplished by a number of sensors throughout the system that either send signals to controller **105** or are queried by controller **105**. For example, with reference to BPMED unit **110**, such monitors may include one or more pH gauges to monitor a pH within the units as well as pressure sensors to monitor. A pressure among the compartments in BPMED unit **110** is to avoid mechanical damage to the BPMED membrane stack and the unwanted mixing of different solution streams within the membrane stack. Other monitors include one or pressure monitors associated with BPMED unit **110** to minimize the expulsion of gases within the unit.

The above-described system may be used to collect CO₂ from seawater or any other liquid source. Such collection may serve to reduce a concentration of CO₂ in the atmosphere and also provide a source of CO₂ may be used in various industries, including, but not limited to, as a fuel source.

In the description above, for the purposes of explanation, numerous specific details have been set forth in order to provide a thorough understanding of the embodiments. It will be apparent however, to one skilled in the art, that one or more other embodiments may be practiced without some of these specific details. The particular embodiments described are not provided to limit the invention but to illustrate it. The scope of the invention is not to be determined by the specific examples provided above but only by the claims below. In other instances, well-known structures, devices, and operations have been shown in block diagram form or without detail in order to avoid obscuring the understanding of the description. Where considered appropriate, reference numerals or terminal portions of reference numerals have been repeated among the figures to indicate corresponding or analogous elements, which may optionally have similar characteristics.

It should also be appreciated that reference throughout this specification to “one embodiment”, “an embodiment”, “one or more embodiments”, or “different embodiments”, for example, means that a particular feature may be included in the practice of the invention. Similarly, it should be appreciated that in the description various features are sometimes grouped together in a single embodiment, figure, or description thereof for the purpose of streamlining the disclosure and aiding in the understanding of various inventive aspects. This method of disclosure, however, is not to be interpreted as reflecting an intention that the invention requires more features than are expressly recited in each claim. Rather, as the following claims reflect, inventive aspects may lie in less than all features of a single disclosed embodiment. Thus, the claims following the Detailed Description are hereby expressly incorporated into this Detailed Description, with each claim standing on its own as a separate embodiment of the invention.

What is claimed is:

1. A method comprising:
 - increasing a concentration of dissolved inorganic carbon in a volume of seawater to produce modified seawater;

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acidifying the modified seawater by flowing the modified seawater into an acidified solution compartment in an electro dialysis unit;

flowing the seawater from a second volume of the seawater into a basified solution compartment in the electro dialysis unit at the same time as flowing the modified seawater into the acidified solution compartment, wherein the second volume of the seawater is unmodified; and

collecting an amount of carbon dioxide from the acidified seawater.

2. The method of claim 1, wherein increasing a concentration of dissolved inorganic carbon in the seawater comprises heating the seawater.

3. The method of claim 2, wherein heating the seawater comprises exposing the seawater to low grade heat.

4. The method of claim 2, wherein heating the seawater comprises exposing the seawater to solar thermal energy.

5. The method of claim 1, wherein prior to acidifying the modified seawater, the method comprises cooling the modified seawater.

6. The method of claim 1, wherein after acidifying the modified seawater, the method comprises transferring the acidified seawater to a tank, and collecting an amount of carbon dioxide from the acidified seawater comprises collecting the carbon dioxide that escapes from the acidified seawater in the tank.

7. The method of claim 5, wherein prior to transferring the acidified seawater to the tank, the method comprises evacuating the tank of air.

8. The method of claim 1, wherein collecting an amount of carbon dioxide comprises vacuum stripping the carbon dioxide from the acidified seawater.

9. A method comprising:

removing a portion of water molecules from a volume of seawater comprising an amount of dissolved inorganic carbon to form modified seawater;

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after removing a portion of water molecules from the volume of seawater, introducing the modified seawater into an acidified solution compartment in an electro dialysis unit;

acidifying the seawater in the acidified solution compartment;

at the same time as introducing the modified seawater into the acidified solution compartment, introducing the seawater from a second volume of the seawater into a basified solution compartment in the electro dialysis unit, wherein the second volume of the seawater is unmodified; and

collecting an amount of carbon dioxide from the acidified seawater.

10. The method of claim 9, wherein removing a portion of water molecules from a volume of seawater comprises heating the seawater.

11. The method of claim 10, wherein heating the seawater comprises exposing the seawater to low grade heat.

12. The method of claim 10, wherein heating the seawater comprises exposing the seawater to solar thermal energy.

13. The method of claim 10, wherein prior to introducing the seawater into an electro dialysis unit, the method comprises cooling the modified seawater.

14. The method of claim 9, wherein after acidifying the modified seawater, the method comprises transferring the acidified seawater to a tank, and collecting an amount of carbon dioxide from the acidified seawater comprises collecting the carbon dioxide that escapes from the acidified seawater in the tank.

15. The method of claim 14, wherein prior to transferring the acidified seawater to the tank, the method comprises evacuating the tank of air.

16. The method of claim 9, wherein collecting an amount of carbon dioxide comprises vacuum stripping the carbon dioxide from the acidified seawater.

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